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1. all matter is composed of atoms, 2.

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atoms of a given element are identical, atoms of different elements differ from each other, 3. atoms cannot be subdivided, created, or destroyed, 4. atoms combine in simple whole-number ratios to form compounds, 5. in chemical reactions elements are combined, separated, or

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1. Atoms are divisible into smaller, subatomic particles. 2. A given element can have atoms with different masses, called isotopes. The

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formation of water according to the equation. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$. shows that 2 molecules (made of 4 atoms) of hydrogen and 1 molecule (made of 2 atoms) of oxygen produce 2 molecules of water.

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isotope has an atomic number 51 and
a mass number 123.

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chemistry unit 3 study guide. matter can be changed from one form into another, mixtures can be separated or made, and pure substances can be decomposed, but the total amount of mass remains constant. If two elements form more than one

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Compound between them, then the ratios of the masses of the second element which combine with a fixed mass of the first element will be ratios of small whole numbers.

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